Physicochemical Study of Spiropyran-terthiophene Derivatives: Photochemistry and Thermodynamics

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The photochemistry and thermodynamics of two terthiophene (TTth) derivatives bearing benzospiropyran (BSP) moieties, 1-(3,3''-dimethylindoline-6'-nitrobenzospiropyranyl)-2-ethyl 4,4''-didecyloxy-2,2',5',2''-terthiophene-3'-acetate (BSP-2) and 1-(3,3''-dimethylindoline-6'-nitrobenzospiropyranyl)-2-ethyl 4,4''-didecyloxy-2,2',5',2''-terthiophene-3'-carboxylate (BSP-3), differing only by a single methylene spacer unit, have been studied. The kinetics of photogeneration of the equivalent merocyanine (MC) isomers (MC-2 and MC-3, respectively), the isomerisation properties of MC-2 and MC-3, and the thermodynamic parameters have been studied in acetonitrile, and compared to the parent, non-TTth-functionalised, benzospiropyran derivative, BSP-1. Despite the close structural similarity of BSP-2 and BSP-3, their photochemical properties were found to differ significantly; examples include activation energies ($E_a$) = 75.05 kJ mol$^{-1}$, $E_a$ = 100.39 kJ mol$^{-1}$ and entropies of activation ($\Delta S_v$) = -43.38 J K$^{-1}$ mol$^{-1}$, $\Delta S_v$ = -37.78 J K$^{-1}$ mol$^{-1}$) for the thermal relaxation from MC to BSP, with the MC-3 value much closer to the unmodified MC-1 value (46.48 J K$^{-1}$ mol$^{-1}$) for this latter quantity. The thermal relaxation kinetics and solvatochromic behaviour of the derivatives in a range of solvents of differing polarity (ethanol, dichloromethane, acetone, toluene and diethyl ether) are also presented.

Differences in the estimated values of these thermodynamic and kinetic parameters are discussed with reference to the molecular structure of the derivatives.

Introduction

Nitrobenzospiropyran (BSP) is a member of a family of photo/thermochromic molecules that have been widely studied due to their ability to be switched reversibly under various external stimuli between two molecular states that exhibit dramatically different properties. This behaviour has attracted great interest in terms of possible applications such as optical memory, biomaterials for targeted therapeutic agents, components for ophthalmic lenses, molecular functionalization of micro channels and as optical probes for the analysis of nanostructures in ionic liquids (ILs).

Photochromism is a phenomenon that occurs when electromagnetic radiation is absorbed by a molecule that undergoes a reversible rearrangement between two forms, such as 1-(2-hydroxyethyl)-3,3-dimethylindoline-6'-nitrobenzospiropyran (BSP-1) and its equivalent merocyanine form (MC-1), with dramatic accompanying changes in the absorption spectra in the visible region.

Benzospiropyran is a molecule that exists as two heterocyclic functional units connected by a tetrahedral sp3 hybridised carbon atom. This atom co-joins two orthogonal planes comprising an indoline subunit and a benzopyran subunit, the common component for all the spiropyran moieties.

Photo-induced isomerisation processes are well known in the literature. In spiropyran, heterocyclic ring cleavage occurs at the $\text{C}_\text{spiro}$-O bond when exposed to UV light, leading to the formation of the merocyanine (MC) species that can be detected via a striking colour change due to the strong absorbance of the MC isomer in the visible region (Figure 1). Upon irradiation with white light, the system reverts back to the original BSP form. The MC isomer absorbance is sensitive to its immediate molecular environment, leading to solvent dependant colour changes, and the solvent can also mediate the equilibrium between the isomers.

The solvatochromism of various benzospiropyran derivatives has been reported previously including nitro- and methoxy-substituted spiro-moieties. Recently, the well-known photo-, thermo- and solvato-chromatic properties of benzospiropyran (BSP) and spiroxazine derivatives have been used by our group to provide new insights into the nanostructure of ILs through in-situ mediation of their switching behaviour. Furthermore the presence of the phenolate anion in the MC isomer provides a weak binding capacity for certain metal ions (d- and f- elements), and the resulting complexes generate new absorbance bands in the visible region, which are directly influenced by the type of metal ion present.

Polythiophenes (PThs) are a class of organic polymeric semiconductors with extremely interesting conductivity, electrochromism and electroluminescence properties, which have been explored for applications such as electrochemical sensors and energy harvesting devices. From a materials chemistry perspective, they are particularly attractive due to their high chemical stability and relative ease of isolation. Other interesting pTh properties previously reported in the literature include thermochromism, ionochromism, photochromism and bio-chromism. Chromatic variations are generated by

![BSP-1 and MC-1](image)

Fig. 1 Molecular structure and isomerisation of 1-(2-hydroxyethyl)-3,3-dimethylindoline-6'-nitrobenzospiropyran (BSP-1) spiropyran.

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concomitant changes or folding of the polymer chains, with consequent dramatic variation in the conjugation length of the whole system, which can be monitored using spectroscopy. Substituents connected to the three-aromatic rings backbone can have a considerable effect on the overall conjugation of the system, influencing the properties and the distribution of the π-electrons in the structure.

As part of our ongoing interest in functionalised PThs, we recently reported the first synthesis of a BSP-functionalised poly(terthiophene) and undertook a detailed study of its complex electrochemical and photochemical properties. The BSP was attached to the terthiophene monomer (acTTh, Figure 2) by an acetate linkage (BSP-2, Figure 2). Given the multiple effects of the spiropryan on the electrochemical behaviour of the poly(terthiophene), we were interested in how these might vary depending on the distance of the spiropryan from the polymer backbone. Therefore, we synthesised the BSP-terthiophene monomer in which the ester linkage was directly attached to the terthiophene moiety (BSP-3, Figure 2), differing from BSP-2 only by one carbon atom. Prior to studying its electrochemical polymerisation and resulting polymer properties, we deemed it important to determine the photochemical properties of BSP-3 and compare them to BSP-1 and BSP-2.

Therefore, in this work, the photochemical behaviour of the two BSP-functionalised terthiophene monomers, 1-(3,3′′-dimethylindolin-6′-nitrobenzospiropyranyl)-2-ethyl 4,4′′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-carboxylate (BSP-2) and 1-(3,3′′-dimethylindolin-6′-nitrobenzospiropyranyl)-2-ethyl 4,4′′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-acetate (BSP-1) and compare them to important to determine the photochemical properties of the terthiophene moiety (monomer in which the ester linkage was directly attached to the terthiophene moiety). Depending on the distance of the spiropyran from the polymer backbone, there is a considerable effect on the overall conjugation of the system, influencing the properties and the distribution of the π-electrons in the structure.

Fig. 2 Structures of the TTh and TTh-BSP derivatives studied: Methyl 4,4′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-acetate (AcTTh), 4,4′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-carboxylic acid (cbTTh), 1-(3,3′′-dimethylindolin-6′-nitrobenzospiropyranyl)-2-ethyl 4,4′′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-acetate (BSP-2) and 1-(3,3′′-dimethylindolin-6′-nitrobenzospiropyranyl)-2-ethyl 4,4′′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-carboxylate (BSP-3).

Experimental

Materials

1-(2-Hydroxyethyl)-3,3-dimethylindolin-6′-nitrobenzospiropyran (BSP-1) was purchased from TCI chemicals, Ireland, and was used without further purification. Ethyl (3,3′′-dimethylindolin-6′-nitrobenzospiropyranyl)-4,4′′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-acetate (BSP-2) was synthesised as described previously. Ethyl (3,3′′-dimethylindolin-6′-nitrobenzospiropyranyl)-4,4′′-didecyloxy-2,2′:5′,2′′-terthiophene-3′-carboxylate (BSP-3) was synthesised as follows: 4,4′′-Didecyloxy-2,2′:5′,2′′-terthiophene-3′-carboxylic acid (cbTTh) (0.083 g, 0.138 mmol) and diethylxycarbodiimide (DCC) (0.03 g, 0.143 mmol) and 4-dimethylaminopyridine (0.0013 g, 0.011 mmol) were dissolved in DCM (12 ml) at 0 °C. 1-(2-Hydroxyethyl)-3,3′′-dimethylindolin-6′-nitrobenzospiropyran (BSP-1) (0.04 g, 0.11 mmol) in DCM (10 ml) was added dropwise over 30 min to the stirring solution. After BSP-1 was completely added, the reaction mixture was allowed to reach 20 °C. After a further 90 min, 50% of the DCM was removed under reduced pressure, cold ether (10 ml) was added to precipitate the urea salt of DCC. This precipitate was removed by filtration and the filtrate was reduced to give a brown oil, which was purified on a silica column, eluting with DCM. The resulting product was a yellowish oil, which solidified on standing (0.081 g, 77%). After purification via column chromatography, the BSP-3 was stored in the dark in a sealed glass vial under nitrogen. Characterisation data are available in the ESI section.

The solvents used for the photo- and thermo-chronic analyses (acetonitrile (ACN), ethanol (EtOH), toluene (Tol), acetone (Acet), dichloromethane (DCM) and diethyl ether (DetEth)) were purchased from TCI chemicals, Ireland. Purified acetonitrile (ACN) was used to ensure equilibrium in the solvent was reached. Diluted solutions (10^{-7} M) of BSP-1, BSP-2 and BSP-3 were exposed to UV light and the spectrum of each was taken at intervals from t = 1 minute to t = 45 minutes.

A single exponential model (equation (1)) was used to determine the ring opening rate constants for BSP-1, BSP-2 and BSP-3.

\[ \chi = a \times (1 - e^{-(kt)}) + b \]  

(1)
where $y$ is the absorbance at $\lambda_{\text{max}}$ (assumed to be proportional to the concentration of the merocyanine isomer), $a$ is a scaling factor, $k$ is the first order rate constant (s$^{-1}$), $b$ is the baseline offset and $t$ is time (s).

Absorption spectra were recorded using a Perkin Elmer Lambda 900 UV-Vis/NIR spectrometer equipped with a thermostatted temperature control system (PTP-1 Peltier System). Photo-conversion of BSP to MC was achieved using an in-house fabricated UV light source consisting of three UV LEDs (Roithner LaserTechnik GmbH, emission $\lambda_{\text{max}} = 375$ nm). The LEDs were powered from a 3.75 V current source, which provided an optical power output of 6 mW/cm$^2$ from each LED. The white light radiation was provided via a DDL 150 W source obtained from Polytec GmbH Waldbronn.

To determine the thermodynamic parameters of the samples and obtain an estimation of the transition state equilibrium of activation ($K^\dagger$) for the $\text{MC} \rightleftharpoons \text{BSP}$ equilibrium, standard solutions of BSP-1, BSP-2 and BSP-3 were prepared in concentrations up to $10^{-6}$ M in acetonitrile. The solutions were stored in volumetric flasks in the dark under nitrogen before the absorbance measurements were acquired. Thermal relaxation analysis of the thermodynamically favoured reversion of MC to BSP was measured at $5^\circ$C increments from 283 to 308 K over a period of 1 hour. Samples were irradiated in the dark with the UV-VIS LED system in a sealed quartz cuvette at room temperature for 3 minutes prior to acquisition of the spectra. The study of the $\text{BSP} \rightleftharpoons \text{MC}$ equilibrium was performed using the LED system in the spectrometer chamber in order to minimise any influence of external light. The $\text{MC} \rightleftharpoons \text{BSP}$ thermal relaxation typically follows first order kinetics in molecular solvents and the rate is not sensitive to oxygen, but is strongly medium dependent.

First-order rate constants for the thermal relaxation were determined in this work using Microsoft Excel Solver, by fitting the absorbance data to equation (2), where $y$ is the absorbance value at the $\lambda_{\text{max}}$, $a$ is the absorbance at $t = 0$, $k$ is the rate constant, $t$ is time and $b$ is a baseline offset, as for equation (1)$^{11}$.\[ y = ae^{-kt} + b \number{(2)} \]The Arrhenius equation (3) was used to plot the linear temperature dependence of the rate constant for $\text{MC} \rightleftharpoons \text{BSP}$ thermal relaxation and find the activation energy ($E_a$) and the pre-exponential factor $A$.

\[ \ln k = \frac{E_a}{RT} + \ln A \number{(3)} \]where $k$ = rate constant (s$^{-1}$), $E_a$ = Activation energy (kJ mol$^{-1}$), $R$ = gas constant (J mol$^{-1}$ K$^{-1}$), $T$ = temperature (K) and $A$ = pre exponential factor (s$^{-1}$)

Following the generalised $\text{BSP} \rightleftharpoons \text{MC}$ system (Figure 1), the Eyring equation (4) was used to determine the thermodynamic properties at the transition state equilibrium of activation ($K^\dagger$), entropy of activation ($\Delta S^\dagger$), and the enthalpy of activation ($\Delta H^\dagger$). A derivative of this equation, equation (5), can be used to derive the equilibrium of activation ($K^\dagger$)$^{16, 17, 22}$.

\[ \ln(k / T) = -\Delta H^\dagger / RT + \ln(k / h) + \Delta S^\dagger / R \number{(4)} \]where $k$ = rate constant of the forward reaction, $T$ = temperature (K), $\Delta H^\dagger$ = enthalpy of activation, $\Delta S^\dagger$ = entropy of activation, $R$ = gas constant, $k_B$ = Boltzmann constant and $h$ = Planck’s constant.

Thermodynamic activation properties such as $\Delta S^\dagger$, $\Delta H^\dagger$ and $K^\dagger$, for each $\text{MC}$ derivative were calculated using the Eyring equation. Equation (6) was used to calculate the Gibbs free energy of activation ($\Delta G^\dagger$).

\[ \Delta G^\dagger = \Delta H^\dagger - T \Delta S^\dagger \number{(6)} \]where $\Delta G^\dagger$ = Gibbs free energy of activation (kJ mol$^{-1}$), $\Delta H^\dagger$ = enthalpy of transition state (kJ mol$^{-1}$), $T$ = temperature (K) and $\Delta S^\dagger$ = entropy of the transition state (J K$^{-1}$ mol$^{-1}$).

The solvatochromic study was performed at room temperature in 6 different solvent systems (ethanol, toluene, acetone, acetonitrile, diethyl ether and dichloromethane). The solutions (all $10^{-4}$ M) were sealed and stored under nitrogen overnight before the analysis was performed. Spectra were taken under controlled temperature conditions using the LED system described above to illuminate the samples.

Fluorescence (emission) spectra were recorded with a Perkin Elmer Luminescence Spectrometer LS 50B. Solutions diluted up to $10^{-3}$ M in acetonitrile of the two freestanding TTh molecules, BSP-1, BSP-2 and BSP-3 were made. Excitation wavelengths were set as follows: 358 nm (acTTh), 362 nm (cbTTh), 360 nm (BSP-2 and BSP-3). The solutions were stored in sealed volumetric flasks for 24 hrs under nitrogen in the dark. Photo-isomerisation to the relatives MC-1, MC-2 and MC-3 was carried out with in situ LED illumination (537 nm) for 180 seconds.

Results and Discussion

Material synthesis

The synthesis of BSP-3 in 77% yield by a base-catalyzed substitution of BSP-1 onto cbTTh was straightforward and its spectral data (see Electronic Supplementary Information (ESI)) was consistent with this structure. As expected, its $^1$H NMR spectrum was near identical to that of BSP-2 except for the absence of the methylene signal at $\delta = 7.3$.

Photochromic properties

It is well known that $\text{MC}$ derivatives typically exhibit negative solvatochromism, meaning that the absorption maximum undergoes a hypsochromic (blue) shift as the solvent polarity increases, due to interactions between the solute and the solvent molecules that increase the energy gap between the $\text{MC}$ ground and the excited states.$^{18-21}$ The focus of this study is whether the spectroscopic properties and switching behaviour of the BSP group is affected by attachment of the differing TTh sub-units.
Derivatives in ACN. For the formation of BSP giving an order in terms of speed of ring opening of equivalent rate constant was higher again (k = 4.45 x 10^-4 M^-1 s^-1) occurs ~1.5 times faster for BSP-2 compared to BSP-3 (k_{BSP,2} = 4.45 x 10^{-4} s^{-1} versus k_{BSP,3} = 2.85 x 10^{-4} s^{-1}). For BSP-1 the equivalent rate constant was higher again (k_{BSP,1} = 1.09 x 10^{-4} s^{-1}), giving an order in terms of speed of ring opening of BSP-1 > BSP-2 > BSP-3.

Extrapolation of the fitted curves using equation (1) allowed estimation of the steady state absorbance values for the derivatives in ACN. For the formation of MC-2, the model suggests the steady state (ca. 0.4 a.u.) is reached in about 150 minutes while for the formation of MC-3 steady state absorbance is relatively constant at ca. 0.50 a.u. after about 200 min. This would be expected due to BSP-2 having a faster ring-opening rate than BSP-3 (Figure 4 (c)). Clearly, there is a significant difference in the steady state absorbance of BSP-1 relative to BSP-2 and BSP-3: the detected maximum absorbance for BSP-1 after 45 minutes of exposure to the LED device was 1.84 a.u. at 563 nm versus 0.31 a.u. at 568 nm and 0.39 a.u. at 572 nm for BSP-2 and BSP-3, respectively. This is most likely due to competing absorbance by the TTh moiety present in BSP-2 and BSP-3, which reduces the availability of UV-photons for driving the ring-opening process.

UV-visible spectra of the two spiropyran-functionalised terthiophenes show two major absorbance bands with maxima at 352 nm and 362 nm for BSP-2 and BSP-3, respectively (TTh absorbance, BSP contribution not shown) and 568 nm and 572 nm for MC-2 and MC-3, respectively (MC absorbance), as shown in Figure 3 (a) and (b). All solutions were kept in the dark for 24 hr at 293 K to establish thermodynamic equilibrium prior to UV illumination. In Figure 3 (a), the absorbance maximum at ca. 360 nm is greater for MC-2 compared to BSP-2, whereas they are the same for BSP-3 and MC-3 (Figure 3 (b)). This suggests that there is a change in the electronic structure of MC-2 relative to BSP-2, which does not occur with BSP-3/MC-3.

**Kinetics of photo-induced ring opening**

Figures 4 (a) and (b) show the spectra used to obtain the λ_{max} values for substitution into the model (equation 1), while Figure 4 (c) gives the growth in absorbance over time (293 K) at the λ_{max} for MC-1, MC-2 and MC-3 (spectra for MC-1 are given in ESI, Figure S1). The ring opening process analysed using equation (1) occurs ~1.5 times faster for BSP-2 compared to BSP-3 (k_{BSP,2} = 4.45 x 10^{-4} s^{-1} versus k_{BSP,3} = 2.85 x 10^{-4} s^{-1}). For BSP-1 the equivalent rate constant was higher again (k_{BSP,1} = 1.09 x 10^{-4} s^{-1}), giving an order in terms of speed of ring opening of BSP-1 > BSP-2 > BSP-3.

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**Figure 3** Baseline corrected UV-visible spectra before and after illumination of 10^{-5} M solutions of (a) BSP-2 with its MC-2 absorption band at 568 nm and (b) BSP-3 with its MC-3 absorption band at 572 nm, using the 3 UV LED light source with λ_{max} emission = 375 nm; exposure time = 3 minutes.

**Figure 4** Exposure effects to 375 nm UV LEDs: (a) Baseline corrected conversion of BSP-2 (10^{-5} M in ACN) to the isomer MC-2. (b) Baseline corrected conversion of spiropyran BSP-3 (10^{-6} M in ACN) to the isomer MC-3. (c) Exponential curves extrapolated to steady state for generation of MC isomers versus time of UV irradiation with experimental results superimposed for MC-2 and MC-3. Values were taken at λ_{max} of MC-1 = 563 nm, MC-2 = 568 nm, MC-3 = 572 nm.

Further indications supporting this statement might be found by
comparing the ring opening of BSP-2 and BSP-3. Extrapolation of the curves (Figure 4 (c)) suggests that the process reaches steady state after ca. 150 minutes for MC-2 and ca. 200 minutes for MC-3. This model predicts final maximum absorbances for $10^{-4}$ M acetonitrile solutions of ca. 0.39 a.u. for MC-2, and ca. 0.5 a.u. for MC-3, considerably smaller if compared to derivatives already reported in literature for BSP-1 (1.8 a.u.).

The calculated rate of ring opening in the described process, BSP→MC, is fastest for BSP-2 compared to BSP-3, but both are slower than BSP-1 by an order of magnitude.

**Kinetics and thermodynamics of ring closure**

Table 1 shows the rate constants for ring closure (MC→BSP) over the temperature range 283 K – 308 K for BSP-1, BSP-2 and BSP-3, and the ratios of the rate constants for each form. The MC-3/MC-2 ratio steadily increases from 0.5 at 283 K to 1.4 at 308 K, meaning that at 283 K, MC-2 relaxation is faster, whereas at 308 K, MC-3 is faster. This pattern of behaviour is repeated for MC-1 compared to MC-2, with the ratio = 0.42 at 283 K, increasing to 1.15 at 308 K. In contrast, the MC-1/MC-3 ratio remains relatively constant over the temperature range studied, at about 0.8. These results suggest that the temperature affects ring closure kinetics in a similar manner for BSP-2 compared to the parent BSP-1 (increases around 3-fold), but a different type of dependency is observed with BSP-2 (remains relatively constant). Furthermore, at 283 K MC-2 reversion to BSP-2 is the fastest for the three derivatives studied, whereas at 308 K, it is the slowest.

**Table 1** Rate constants $k$ and their ratios for MC→BSP thermal relaxation for MC-1, MC-2 and MC-3 ($10^{-4}$ M in ACN) over 283 – 308 K averaged from 3 replicates. Standard deviations reported in brackets.

<table>
<thead>
<tr>
<th></th>
<th>$k$ (10$^{-5}$ s$^{-1}$) in ACN</th>
<th>$k$ (10$^{-5}$ s$^{-1}$)</th>
<th>$k$ (10$^{-5}$ s$^{-1}$)</th>
<th>$k$ (10$^{-5}$ s$^{-1}$)</th>
<th>$k$ (10$^{-5}$ s$^{-1}$)</th>
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</thead>
<tbody>
<tr>
<td></td>
<td>283 K</td>
<td>288 K</td>
<td>293 K</td>
<td>298 K</td>
<td>303 K</td>
</tr>
<tr>
<td>MC-1</td>
<td>563</td>
<td>(0.0002)</td>
<td>1.06</td>
<td>(0.0011)</td>
<td>2.2</td>
</tr>
<tr>
<td>MC-2</td>
<td>568</td>
<td>(0.0002)</td>
<td>1.2</td>
<td>(0.0002)</td>
<td>2.6</td>
</tr>
<tr>
<td>MC-3</td>
<td>572</td>
<td>(0.0004)</td>
<td>0.6</td>
<td>(0.0004)</td>
<td>1.5</td>
</tr>
<tr>
<td>Ratios MC 1/2</td>
<td>0.42</td>
<td>0.40</td>
<td>0.50</td>
<td>0.52</td>
<td>0.77</td>
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<tr>
<td>Ratios MC 1/3</td>
<td>0.83</td>
<td>0.70</td>
<td>0.62</td>
<td>0.76</td>
<td>0.82</td>
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<tr>
<td>Ratios MC 3/2</td>
<td>0.50</td>
<td>0.57</td>
<td>0.69</td>
<td>0.71</td>
<td>0.93</td>
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</table>

Figure 5 shows the absorbance decrease at 568 nm for MC-3 in ACN ($10^{-4}$ M) at different temperatures (similar spectra for MC-2 and MC-1 are given in the ESI, Figures S3 and S4, respectively), from which the rate constants for ring closing (thermal relaxation) at each temperature were obtained by fitting a first order rate equation.

The resulting Arrhenius and Eyring plots are shown in Figure 6 (a) and (b). From the Arrhenius plots, the activation energy ($E_a$) and the pre-exponential factor were derived (Table 2) for the thermal relaxation process. Using Eyring’s transition state theory (Equation (4)), the activation thermodynamic parameters, $\Delta S^*$ (entropy of activation), $\Delta H^*$ (enthalpy of activation) and $\Delta G^*$ (Gibbs free energy of activation), for the thermal relaxation process can be estimated, (Table 2, Figure 6 (b)).

![Fig. 5](https://example.com/figure5.png)
merocyanine isomer is present, and this would favour the adoption of a more ordered configuration, which is consistent with the negative value of $\Delta S^\ddagger$ for BSP-2. This explanation is also supported by the difference in $\Delta H^\ddagger$ between the three spiropyrans: $\Delta H^\ddagger$ for BSP-2 is ~20 kJ mol$^{-1}$ lower than BSP-1 and BSP-3. This suggests that the MC-2 adopts a ground state conformation somewhat closer to the transition state leading to the spiropyran (BSP-2) than MC-3 does on isomerising to BSP-3. Given that the formation of spiropyran from merocyanine typically requires rotation about the three bonds linking the indoline and phenolate components of the merocyanine, as discussed extensively by Wojtyk et al., a $\pi$-$\pi$ interaction between the terthiophene and the phenolate component of MC-2 that brought the phenolate oxygen (in polar solvents) or quinone oxygen (in non-polar solvents) closer to the indoline nitrogen should lower $\Delta H^\ddagger$. Such an MC conformation could be determined by $\pi$-$\pi$ interactions during the ring opening of the spiropyran to form the merocyanine initially.

Table 2: Thermodynamic parameters expressed as a mean of 3 replicates and standard deviation (standard deviations values are reported in brackets) for MC-2 $\leftrightarrow$ BSP-2 and MC-3 $\leftrightarrow$ BSP-3 thermal relaxation compared to the results obtained from a reference solution of BSP-1 $10^{-3}$ M in ACN.

<table>
<thead>
<tr>
<th>MC</th>
<th>$\lambda_{max}$ (nm)</th>
<th>$k_{293}$ ($10^{-3}$s$^{-1}$)</th>
<th>$E_a$ (kJ mol$^{-1}$)</th>
<th>$A$ (s$^{-1}$)</th>
<th>$\Delta S^\ddagger$ (JK mol$^{-1}$ K$^{-1}$)</th>
<th>$\Delta G^\ddagger$ (kJ mol$^{-1}$)</th>
<th>$\Delta H^\ddagger$ (kJ mol$^{-1}$)</th>
<th>$K^\ddagger$ (10$^{-3}$)</th>
</tr>
</thead>
<tbody>
<tr>
<td>MC-1</td>
<td>563</td>
<td>2.2</td>
<td>101.49 (±2.06)</td>
<td>3.51 x 10$^{14}$ (±2.83)</td>
<td>46.48 (±2.92)</td>
<td>86.59 (±0.082)</td>
<td>99.04 (±2.06)</td>
<td>3.56 (±0.108)</td>
</tr>
<tr>
<td>MC-2</td>
<td>568</td>
<td>4.3</td>
<td>75.05 (±1.5)</td>
<td>1.35 x 10$^{11}$ (±0.9)</td>
<td>43.38 (±2.07)</td>
<td>85.17 (±0.18)</td>
<td>73.05 (±1.5)</td>
<td>6.57 (±0.73)</td>
</tr>
<tr>
<td>MC-3</td>
<td>572</td>
<td>3.6</td>
<td>100.39 (±3.06)</td>
<td>8.62 x 10$^{13}$ (±0.24)</td>
<td>37.78 (±2.81)</td>
<td>85.94 (±0.23)</td>
<td>96.74 (±1.42)</td>
<td>5.93 (±0.74)</td>
</tr>
</tbody>
</table>

Solvatochromic effects

Samples of BSP-2 and BSP-3 were dissolved in selected solvents and kept in the dark for 24 hr before solvatochromic studies were carried out. For this study, a polar protic solvent (ethanol) and polar aprotic solvents (acetone, acetonitrile) were used and their effect compared to non-polar solvents (toluene, dichloromethane and diethyl ether). Figure 7 shows spectra obtained for MC-3 in a number of solvents after irradiation for 180 seconds with UV light (375 nm), along with inset photographs of the resulting solutions. In addition, the thermal relaxation rate constants were estimated at 293 K (Figures 8 (a) and (b), extrapolated from Figure S2 (a) and (b), ESI) in each solvent. The formation of J-aggregates (Figure 7) in non-polar solvents typical of many BSP derivatives is evident from the appearance of new absorbance bands at 605 nm and 599 nm for MC-3, and at 605 nm and 597 nm for MC-2 (Figure S5, ESI), in toluene and diethyl ether, respectively.

Table 3 shows clearly that for both MC-2 and MC-3, the $\lambda_{max}$ is blue-shifted as the solvent polarity increases (negative solvatochromism) from toluene through to ethanol (605 nm to 552 nm for MC-2, 552 nm to 557 nm for MC-3). As is well established in the literature, merocyanines exhibit negative solvatochromism as a result of both a decrease in the dipole moment on electronic excitation and an increase in the dipolar nature of the ground state. It has also been shown that this tends...
to be independent of the nature of the N-substituent, and this is indeed the case for MC-2 and MC-3, which show near identical solvatochromic behaviour.

In contrast, the rate constant for ring closure (k) decreases by an order of magnitude from $24.3 \times 10^3 \text{ s}^{-1}$ to $2.4 \times 10^2 \text{ s}^{-1}$ for MC-2 for increasing solvent polarity, while for MC-3 an even greater rate decrease occurs ($25.7 \times 10^3 \text{ s}^{-1}$ to $0.7 \times 10^3 \text{ s}^{-1}$). The values for the rate constants for ring closure are similar in acetonitrile, acetone, diethyl ether and toluene, implying that solvation of the MC isomers must be quite similar. However, in ethanol this is not the case, as ring closure is four times faster for MC-2.

Table 3: Photo and thermal properties of MC-2 and MC-3 derivatives (10⁻⁴ M) in a selection of solvents. λmax measured after 15 hrs in the dark. The thermal relaxation rate constants were calculated by plotting ln k vs. time, where $k = \text{slope}$.

<table>
<thead>
<tr>
<th>SOLVENT</th>
<th>MC-2</th>
<th>MC-3</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>λmax nm</td>
<td>Abs (k10⁻³ s⁻¹)</td>
</tr>
<tr>
<td>Ethanol</td>
<td>552</td>
<td>0.42</td>
</tr>
<tr>
<td>Acetonitrile</td>
<td>568</td>
<td>0.12</td>
</tr>
<tr>
<td>Acetone</td>
<td>570</td>
<td>0.44</td>
</tr>
<tr>
<td>Dichloromethane</td>
<td>587</td>
<td>0.13</td>
</tr>
<tr>
<td>Diethyl ether</td>
<td>597</td>
<td>0.29</td>
</tr>
<tr>
<td>Toluene</td>
<td>605</td>
<td>0.42</td>
</tr>
</tbody>
</table>

This implies that, even though the dipolar nature of the MC-2 ground state is similar to that of MC-3 (similar λmax values), the ethanol is acting like a less polar solvent with regard to ring closure rate. If intramolecular p-p interactions do indeed control the ground state conformation of MC-2, then MC-2 is in effect partly solvated by the p bonds of terthiophene, a less polar “solvant”, diminishing the full effect of the polar ethanol solvent. Somewhat more surprising is the contrary effect of dichloromethane for which ring closure is almost twice as fast for MC-3 relative to MC-2, even though they appear to be electronically similar in this solvent. A more detailed study needs to be undertaken to probe the reasons for this behaviour.

Fluorescence spectroscopy

The presence of intramolecular p-π interactions in MC-2 would be expected to have a significant impact on the emissive properties of the molecule. The emissions of the terthiophene components of both BSP-2 and BSP-3 and their corresponding merocyanines were investigated since excitation of the molecules at 360 nm would lead to typical terthiophene emissions around 450 nm that overlapped with the edge of the merocyanine absorptions (Figure 3). The corresponding emission spectra, as well as those of BSP-1/MC-1 and the terthiophene acids are presented in Figure 9 (a) and (b), and the relevant data given in Table 4. A comparison of the emission spectra of the two terthiophenes shows, as expected, that acTTh is blue shifted by 28 nm (Figure 9 (a), λmax,em = 439 nm) compared to cbTTh (Figure 9 (b), λmax,em = 467 nm) as a result of the effect of the loss of conjugation of the carboxylic acid. However, both BSP-2 (λmax,em = 464 nm) and MC-2 have notably red-shifted (~25 nm) terthiophene emissions (Figure 9 (a) and Table 4) whilst those of BSP-3 and MC-3 are largely unchanged (Figure 9 (b) and Table 4).
Table 4 Parameters derived from the analysis of the emission spectra of acTTh, chTTh, BSP-2 and BSP-3. Stokes shifts are also indicated.

<table>
<thead>
<tr>
<th>Molecules</th>
<th>λ&lt;sub&gt;exc&lt;/sub&gt; (nm)</th>
<th>λ&lt;sub&gt;em&lt;/sub&gt; (nm)</th>
<th>Stokes Shift (cm&lt;sup&gt;-1&lt;/sup&gt;)</th>
</tr>
</thead>
<tbody>
<tr>
<td>acTTh</td>
<td>28,090 (356)</td>
<td>338</td>
<td>5363</td>
</tr>
<tr>
<td>chTTh</td>
<td>27,472 (364)</td>
<td>358</td>
<td>5042</td>
</tr>
<tr>
<td>BSP-2</td>
<td>20,090 (356)</td>
<td>358</td>
<td>5363</td>
</tr>
<tr>
<td>BSP-3</td>
<td>27,472 (364)</td>
<td>358</td>
<td>5363</td>
</tr>
</tbody>
</table>

Equally significant in this regard is the change in emission intensity for the BSP-2/MC-2 isomers, which is absent for BSP-3/MC-3. While the quantum yields of the molecules were not measured and therefore, it is difficult to draw any conclusions from the relative intensities of the two sets of molecules, the apparent quenching of the MC-2 terthiophene emission is again consistent with an intramolecular π-π interaction between the planar merocyanine moiety and the terthiophene. This may also explain the increase in the merocyanine absorbance at 355 nm (Figure 3 (a)) when BSP-2 is exposed to UV light, which does not manifest when the same experiment is performed with BSP-3, suggesting that a new interaction occurs when BSP-2 is switched to MC-2.

Therefore, these results suggest that the extra methylene group present in the BSP-2 linker chain provides a degree of additional flexibility that leads to enhanced intramolecular interaction between the MC and TTh units, most likely through alignment of π-electrons, as the molecule folds around the linker chain. A 3-d representation of this intramolecular alignment of the MC and TTh π-electrons is shown in Figure 10 (a) and (b). This was produced by generating the energy minimised MC and TTh sub units in Chem-3D Ultra using standard molecular mechanics and arranging these with the appropriate linker chain, while applying typical constraints for the various bond lengths and angles. This tentatively suggests that is possible for the sub-units to align, with a separation as close as 3.5-3.7Å, which is well within the range for π-stacking. As indicated by the shift in the emission spectra, such an interaction is also likely between the BSP and TTh moieties in BSP-2, although perhaps not as strongly given the non-planar nature of the BSP unit. In the case of BSP-3/MC-3, the shorter linker group, coupled with the extended conjugation of the TTh-π electrons into the carbonyl group prevents this interaction from occurring and the fluorescence and absorbance spectra are not affected in the same manner. A more rigorous ab initio study using molecular orbital theory calculations is currently underway and will be the subject of a separate publication.

This explanation is also consistent with the observed differences in the kinetic and thermodynamic behaviour of the BSP-2 and BSP-3. For example, BSP-3 is similar to BSP-1 in terms of thermal relaxation behaviour. Table 1 shows that the ratio of the rate constants remains virtually constant over the temperature range studied (283 K-308 K), whereas for BSP-2, the ratio increases markedly towards the higher end of the range.

This could be explained in terms of disruption of the MC and TTh π-electron alignment as the system energy increases. Similarly, explanations have been advanced for observed patterns in the kinetics the thermal relaxation of the merocyanine form of spiroadaphthoxazine derivatives.

Conclusions

The photochemical and thermodynamic properties of two derivatives, BSP-2 and BSP-3, incorporating terthiophene units attached to spiropyran via the indoline nitrogen atom and ester linkers that differ by one methylene unit have been investigated. In BSP-2, the pyran is separated from the ester group by a methylene group, whereas in BSP-3, the ester group is directly linked to the terthiophene group. This apparently rather small structural difference results in quite striking differences in the characteristics of the derivatives, such as the fluorescence emission spectra (dominated by the terthiophene unit) and the thermal relaxation kinetics of the spiropyran unit. We believe these differences arise from the ability of the π-electron clouds of the thiophene and spiropyran units to align and form intramolecular π-π interactions in BSP-2 due to the flexibility of the longer linker chain. In BSP-3, the terthiophene conjugation extends into the ester group of the linker, and this additional rigidity, coupled with the shorter linker length, prevents a similar type of interaction from occurring. We are now developing a new generation of polymers based on these derivatives that could exploit such effects.

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